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The amount of a substance that contains as many particles as there are atoms in exactly 12 grams of carbon-12. 6.022×10^{23} , the number of atoms or molecules in 1.000 mol. the mass in grams of one mole of a substance. the weighted average of the masses of all naturally occurring isotopes of an element.

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Terms in this set (7) mole The SI base unit used to measure the amount of a substance whose number of particles is the same as the number of atoms of carbon in exactly 12 grams of carbon-12.(p.224)

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Chapter: Chemical Equations and Reactions PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

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Assessment Chapter Test A

Yay, this practice test will determine if you're ready to take the real thing. Questions are taken right out of the book/packet and review lotsa vocab.

Chapter 7 Chem: Practice Test - ProProfs Quiz

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